

CELLS AND FUEL CELLS- SEPARATES

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<p>Cells contain chemicals which react to produce electricity.</p> <p>The voltage produced by a cell is dependent upon a number of factors including the type of electrode and electrolyte.</p> <p>A simple cell can be made by connecting two different metals in contact with an electrolyte.</p> <p>Batteries consist of two or more cells connected together in series to provide a greater voltage.</p> <p>In non-rechargeable cells and batteries the chemical reactions stop when one of the reactants has been used up. Alkaline batteries are non-rechargeable.</p> <p>Rechargeable cells and batteries can be recharged because the chemical reactions are reversed when an external electrical current is supplied.</p>		
<p>Fuel cells are supplied by an external source of fuel (eg hydrogen) and oxygen or air. The fuel is oxidised electrochemically within the fuel cell to produce a potential difference.</p> <p>The overall reaction in a hydrogen fuel cell involves the oxidation of hydrogen to produce water.</p> <p>Hydrogen fuel cells offer a potential alternative to rechargeable cells and batteries.</p>		